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chemical reaction always equals the total mass after a chemical reaction. Law of definite proportion: a given compound always contains exactly the same proportion of elements by mass. For example, water is always 1 g hydrogen for every 8 g oxygen. Law of multiple proportions: When two elements form a series of compounds, the ratios of

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Chemistry, Second Edition Steven S. Zumdahl. D. C. Heath: Lexington, MA, 1989. xxv + 1186 pp. Figs. and tables. 21 X 26 cm. This general chemistry text contains 24 chapters whose topics are arranged in about the same order as the more recently published general chemistry textbooks.

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Chapter 4 (Types of Chemical Reactions and Solution Stoichiometry) - Part 1 Major topics: solution vocabulary, solvation, strong vs. weak electrolytes, molarity, & dilution vocabulary, techniques

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Chapter 11 (Properties of Solutions) Major topics: solution concentration calculations (molarity, percent by mass, mole fraction), steps of solution formation, heat of

Chapter 5 (Gases) - Part 1 Major topics: properties of a gas, pressure, 3 basic gas laws, & combined gas law.

Acid-Base Reactions in Solution: Crash Course Chemistry #8 Last week, Hank talked about how stuff mixes together in solutions. Today, and for the next few weeks, he will talk about the

Chapter 16 (Acid-Base Equilibria) - Part 1 Major topics: Arrhenius vs. Bronsted-Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant (K_a),

Chapter 2 Chemical Principles

Enthalpy: Crash Course Chemistry #18 Energy is like the bestest best friend ever and yet, most of the time we take it for granted. Hank feels bad for our friend and wants

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Chapter 13 (Chemical Equilibrium) - Part 2 Major topics: reaction quotient (Q) & equilibrium calculations (ICE tables)

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Chapter 13 (Chemical Equilibrium) - Part 3 Major topics: equilibrium calculations (ICE tables) cont'd & Le Chatelier's **Principle**.

1. The importance of chemical principles MIT 5.111 **Principles** of **Chemical** Science, Fall 2008 View the complete course: <http://ocw.mit.edu/5-111F08> Instructor: Catherine

Chapter 5 (Gases) - Part 3 & Chapter 13 (Chemical Equilibrium) - Part 1 Major topics: vapor pressure, kinetic molecular theory, diffusion and effusion, equilibrium definitions, & law of mass action (K)

Chapter 12 (Chemical Kinetics) - Part 2 Major topics: method of initial rates cont'd, integrated zero/first/second order rate laws, zero/first/second order reactions

Chapter 2, Zumdahl - Fundamentals Lecture recording from Chapter 2, **Zumdahl** - Fundamentals: Measurements and Calculations.